

**Paper Reference(s)    1SC0/1CH**  
**Pearson Edexcel Level 1/Level 2 GCSE**  
**(9–1)**

**Combined Science**  
**PAPER 2**  
**Higher Tier**

**Diagram Booklet**

**In the boxes below, write your name, centre number and candidate number.**

<b>Surname</b>					
<b>Other names</b>					
<b>Centre Number</b>					
<b>Candidate Number</b>					

## **INSTRUCTIONS**

**There may be spare copies of some diagrams in case you need them.**

**THIS DIAGRAM BOOKLET MUST BE  
RETURNED WITH THE QUESTION PAPER  
AT THE END OF THE EXAMINATION.**

# **Contents**

## **Page**

<b>4</b>	<b>Question 1(a)</b>
<b>5</b>	<b>Question 1(b)(iv)</b>
<b>6</b>	<b>Question 1(b)(iv) (Spare copy)</b>
<b>7</b>	<b>Question 3(c)</b>
<b>8</b>	<b>Question 5(b)</b>
<b>9</b>	<b>Question 5(c)</b>
<b>10</b>	<b>Question 5(d)</b>
<b>11</b>	<b>Question 6(a)(i)</b>
<b>12</b>	<b>Question 6(a)(i) (Spare copy)</b>

Question 1(a)

barium hydroxide	hydrochloric acid	barium chloride	water
solid	aqueous	aqueous	liquid
solid	liquid	solid	aqueous
aqueous	aqueous	solid	liquid
aqueous	liquid	aqueous	aqueous

☐ A

☐ B

☐ C

☐ D

Question 1(b)(iv)

pH of the  
mixture


mass of barium hydroxide in g

**Question 1(b)(iv)**

**pH of the  
mixture**


**mass of barium hydroxide in g**

## Question 3(c)

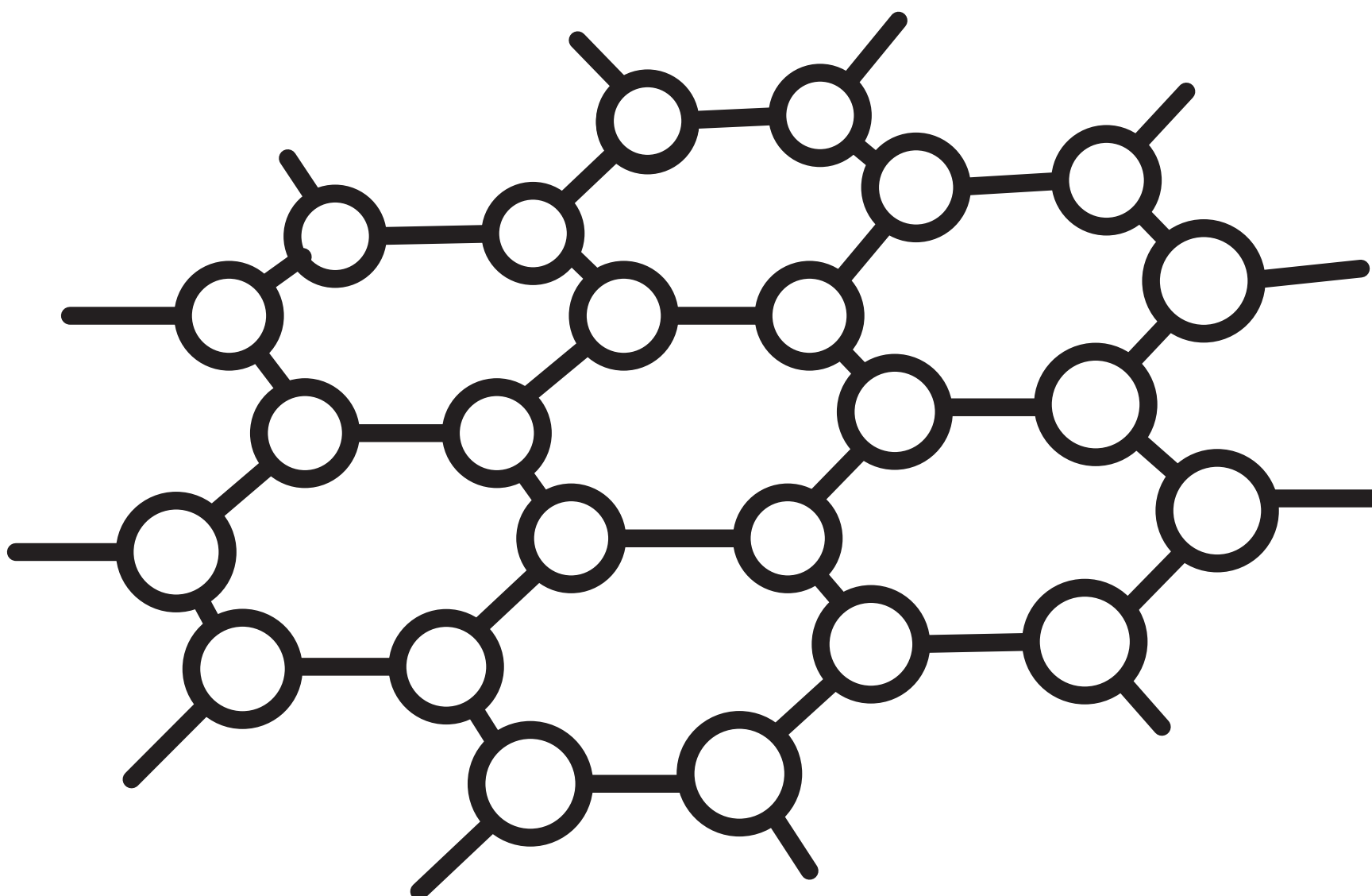
Figure 2

	electrodes	
	anode	cathode
mass of electrode before electrolysis in g	6.43	6.17
mass of electrode after electrolysis in g	5.62	6.95
change in mass in g	−0.81	+0.78

## Question 5(b)

Figure 3

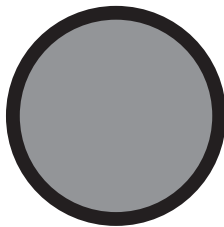
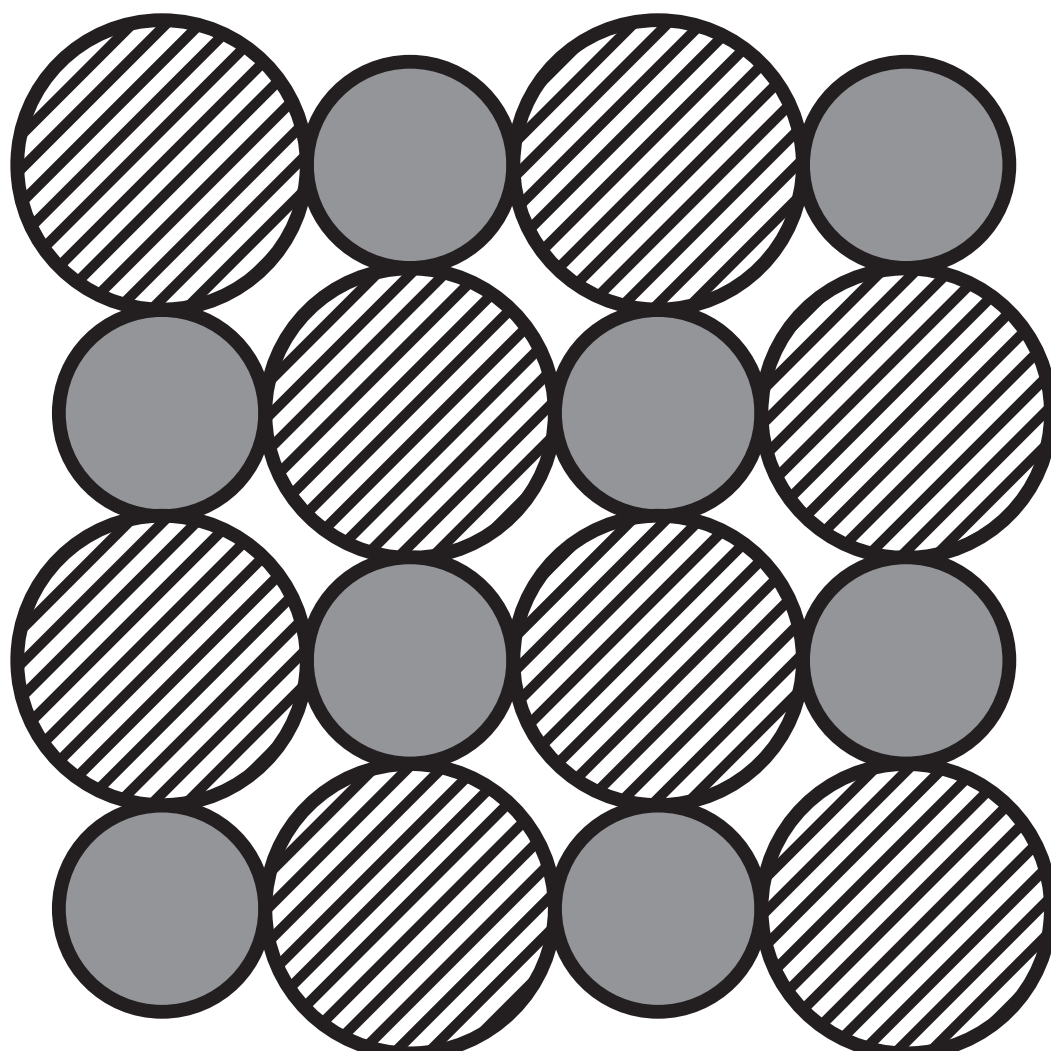
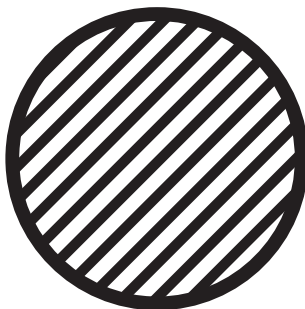
○ carbon atom





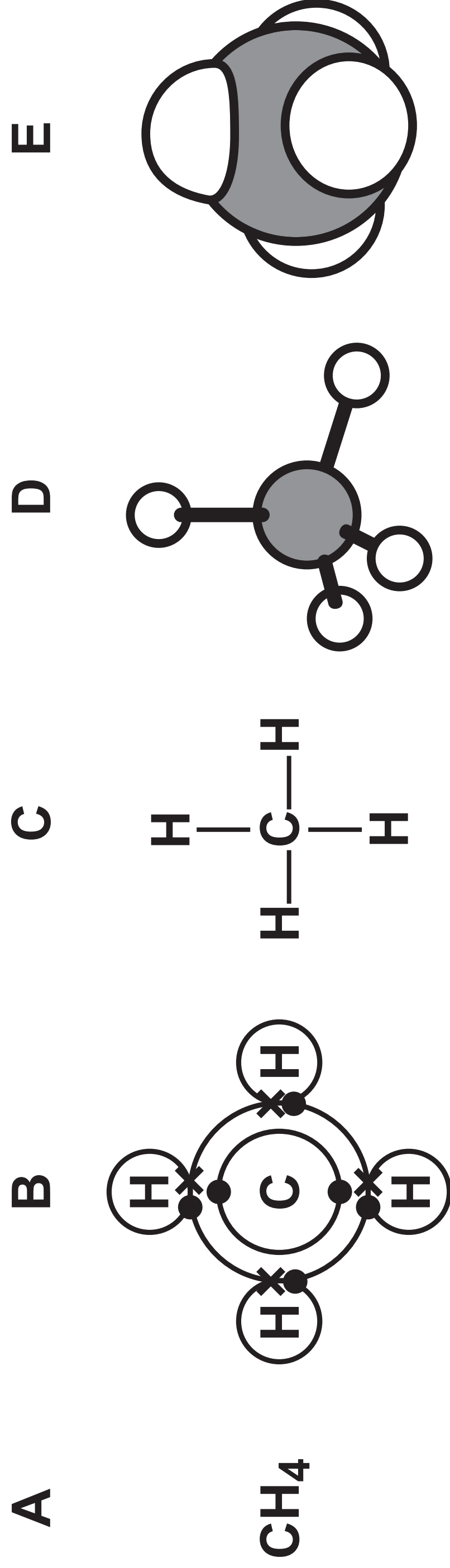
## Question 5(c)

Figure 4

 $K^+$  $Cl^-$ 

# Question 5(d)

Figure 5



Question 6(a)(i)

least reactive		→ most reactive	

Question 6(a)(i)

least reactive

most reactive